



# Cambridge IGCSE™

---

## COMBINED SCIENCE

Paper 1 Multiple Choice (Core)

0653/12

May/June 2021

45 minutes

You must answer on the multiple choice answer sheet.

You will need: Multiple choice answer sheet  
Soft clean eraser  
Soft pencil (type B or HB is recommended)

---

## INSTRUCTIONS

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A, B, C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.

## INFORMATION

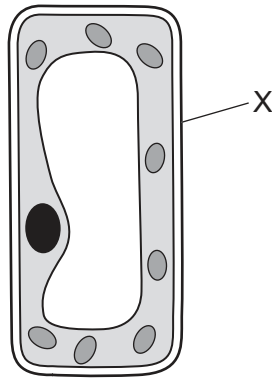
- The total mark for this paper is 40.
- Each correct answer will score one mark.
- Any rough working should be done on this question paper.
- The Periodic Table is printed in the question paper.

---

This document has **16** pages.



- 1 The diagram shows a plant cell as seen under a light microscope.

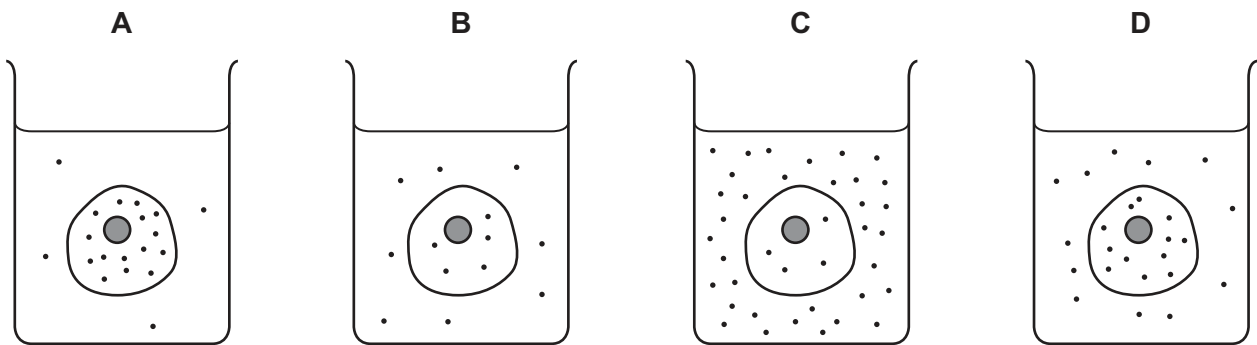


What is the function of the part labelled X?

- A photosynthesis
  - B site of chemical reactions
  - C stores DNA
  - D supports the cell
- 2 The diagrams represent four similar animal cells immersed in blood plasma.

The black dots represent molecules of dissolved oxygen.

Which cell will have oxygen molecules diffusing into it most rapidly?



- 3 A student tests an unknown substance with biuret reagent.

It produces a violet colour.

What is the unknown substance an example of?

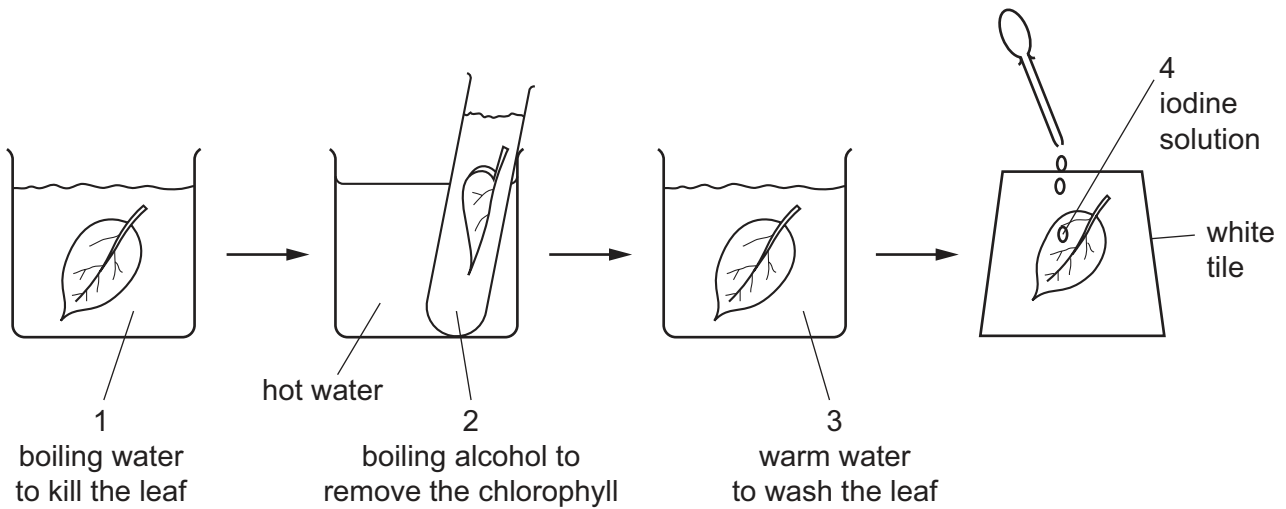
- A fat
- B protein
- C reducing sugar
- D starch

4 Which name is given to biological catalysts?

- A antibodies
- B enzymes
- C hormones
- D platelets

5 The flow diagram shows the stages in testing a green leaf for starch.

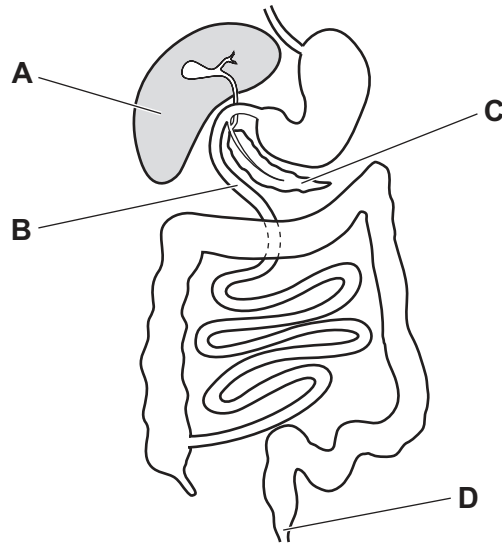
1, 2, 3 and 4 are all liquids.



What are the colours of liquids 2 and 4 for a leaf that contains starch?

	2	4
A	green	blue / black
B	colourless	brown
C	colourless	blue / black
D	green	brown

6 Which part of the alimentary canal carries out digestion **and** absorption?



7 Which statement describes chemical digestion?

- A food particles passing along the alimentary canal
- B large food molecules being broken down into smaller molecules
- C large pieces of food being broken down into smaller pieces
- D nutrients passing through the wall of the small intestine

8 What can be used to test for the presence of carbon dioxide?

- A Benedict's solution
- B ethanol
- C iodine solution
- D limewater

9 Which equation represents aerobic respiration?

- A carbon dioxide + glucose → oxygen + water
- B carbon dioxide + water → glucose + oxygen
- C glucose + oxygen → carbon dioxide + water
- D glucose + water → carbon dioxide + oxygen

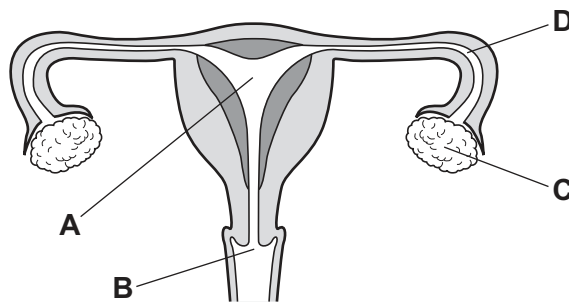
- 10 When the hormone adrenaline is released in humans it causes changes in breathing rate and pupil size.

What are the correct changes?

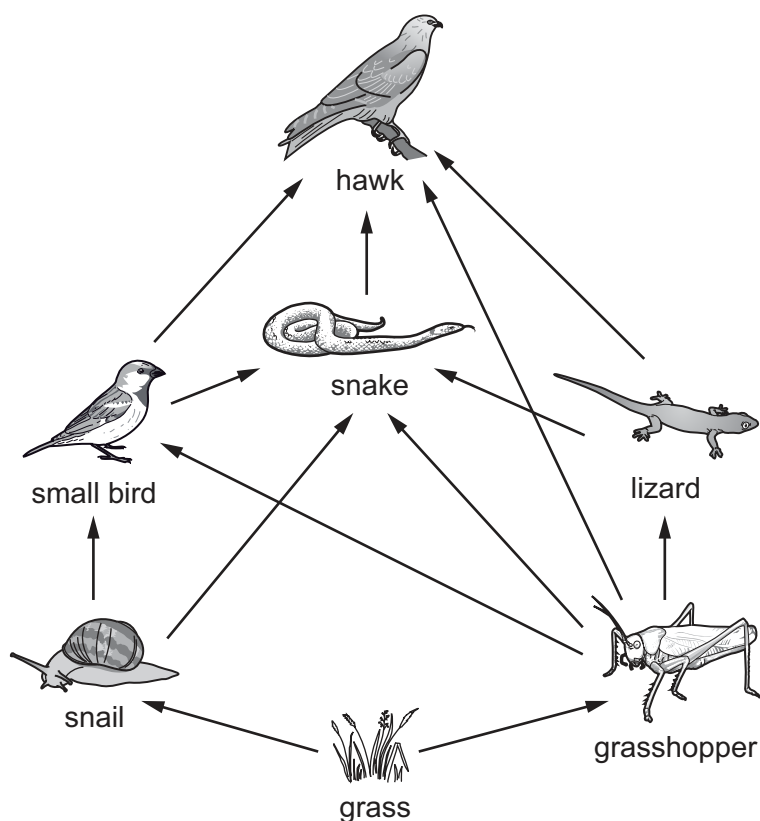
	breathing rate	pupil size
<b>A</b>	decreases	decreases
<b>B</b>	decreases	increases
<b>C</b>	increases	decreases
<b>D</b>	increases	increases

- 11 The diagram shows the human female reproductive system.

Where does fertilisation usually take place?



12 The diagram shows a food web.



Which statement about the snake is correct?

- A It is a consumer and it is a carnivore.
- B It is a producer and it is a carnivore.
- C It is a consumer and it is a herbivore.
- D It is a producer and it is a herbivore.

13 Which process takes carbon dioxide out of the air?

- A combustion
- B decomposition
- C photosynthesis
- D plant respiration

14 What is an example of a physical change?

- A carbon dioxide turning limewater milky
- B the crystallisation of copper(II) sulfate from solution
- C the electrolysis of molten lead(II) bromide
- D the thermal decomposition of calcium carbonate

15 Which substances are mixtures?

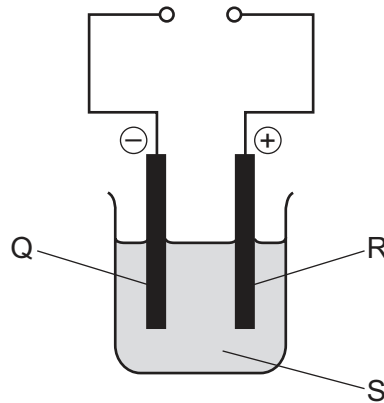
- 1 air
- 2 brass
- 3 sodium chloride

**A** 1 and 2 only    **B** 1 and 3 only    **C** 2 and 3 only    **D** 1, 2 and 3

16 Which statement about the particles is correct?

- A**  ${}^1_1\text{H}$  has the same number of protons as neutrons.
- B**  ${}^2_1\text{H}^+$  has the same number of electrons as neutrons.
- C**  $\text{OH}^-$  contains more protons than electrons.
- D**  $\text{NH}_3$  has the same number of protons as electrons.

17 The apparatus used in an electrolysis experiment is shown.



Which row identifies Q, R and S?

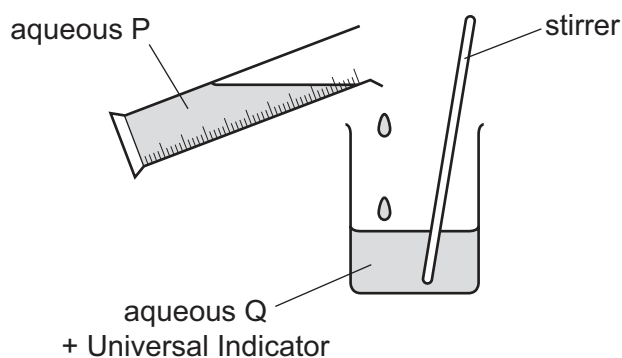
	Q	R	S
<b>A</b>	anode	cathode	electrode
<b>B</b>	cathode	anode	electrode
<b>C</b>	anode	cathode	electrolyte
<b>D</b>	cathode	anode	electrolyte

- 18 Some calcium carbonate and dilute hydrochloric acid start to react. Water is then added to the reaction mixture.

What happens to the rate of the reaction?

- A It decreases.
- B It increases.
- C It stays the same.
- D It stops.

- 19 The diagram shows an experiment to prepare a salt from compounds P and Q.



Aqueous Q has a pH value of 1.

Aqueous P is added until the pH value of the mixture reaches 7.

What are the formulae of compounds P and Q?

	compound P	compound Q
<b>A</b>	HCl	NaOH
<b>B</b>	HNO <sub>3</sub>	H <sub>2</sub> SO <sub>4</sub>
<b>C</b>	KOH	HCl
<b>D</b>	NaOH	NH <sub>3</sub>

- 20 Which two substances form a white precipitate when they are mixed?

- A barium chloride and hydrochloric acid
- B barium chloride and nitric acid
- C silver nitrate and hydrochloric acid
- D silver nitrate and nitric acid



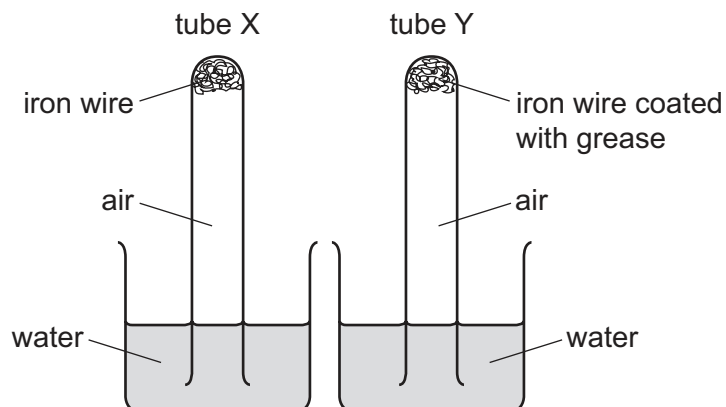
- 21 Which statement describes how the elements change across a period in the Periodic Table from left to right?
- A They change from elements to compounds.
  - B They change from metals to non-metals.
  - C They change from gases to solids.
  - D They change from non-metals to metals.

- 22 Which row shows the properties of a transition element?

	melting point	electrical conductivity	colour of chloride	catalytic properties
A	high	high	white	no
B	high	low	white	no
C	high	high	green	yes
D	low	low	blue	yes

- 23 Which words describe a noble gas?
- A compound, colourless, does not burn in air
  - B element, colourless, burns in air
  - C element, colourless, does not burn in air
  - D element, green, does not burn in air
- 24 Which compound can oxidise carbon?
- A aluminium oxide
  - B copper oxide
  - C magnesium oxide
  - D potassium oxide

25 An experiment is set up to show the effect of air and water on iron.



The experiment is left for one week.

What happens to the water level in each tube?

	tube X	tube Y
<b>A</b>	falls	falls
<b>B</b>	no change	rises
<b>C</b>	rises	rises
<b>D</b>	rises	no change

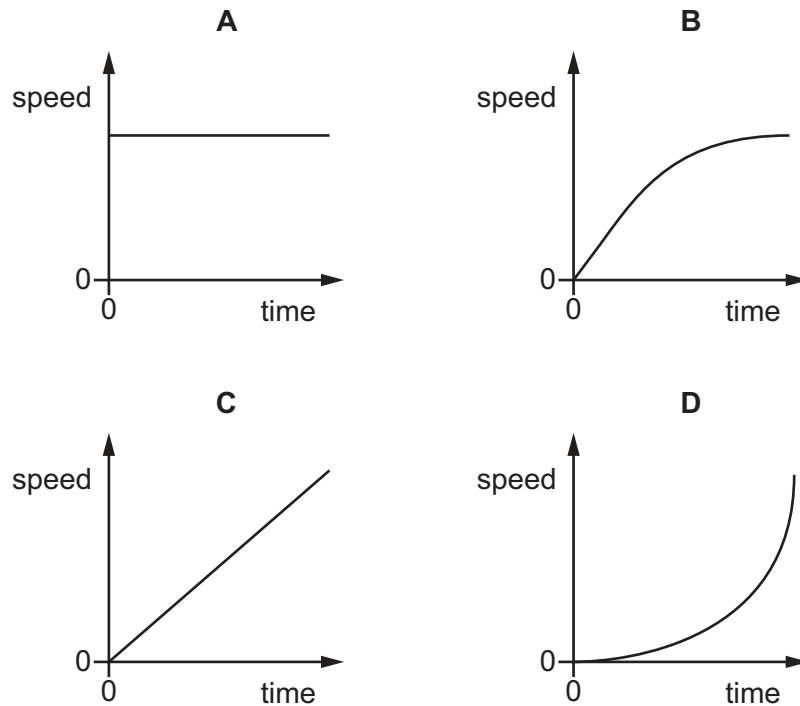
26 Which type of compound contains only carbon and hydrogen?

- A** carbohydrate
- B** carbonate
- C** hydrocarbon
- D** hydroxide

27 Which process produces alkenes?

- A** cracking
- B** fractional distillation
- C** polymerisation
- D** reduction

28 Which speed–time graph represents the motion of an object with constant, non-zero acceleration?



29 The gravitational field strength  $g$  on the surface of the Earth is  $10 \text{ N/kg}$ .

What is the weight of a  $500 \text{ g}$  mass on the surface of the Earth?

- A**  $5.0 \text{ kg}$       **B**  $5.0 \text{ N}$       **C**  $5000 \text{ kg}$       **D**  $5000 \text{ N}$

30 Which row shows apparatus used to measure length, time and volume?

	length	time	volume
<b>A</b>	measuring cylinder	metre rule	stop-clock
<b>B</b>	measuring cylinder	stop-clock	metre rule
<b>C</b>	metre rule	measuring cylinder	stop-clock
<b>D</b>	metre rule	stop-clock	measuring cylinder

31 A block is placed on the ground causing a pressure on the ground.

Which row shows a pair of changes that **must** increase the pressure on the ground?

	weight of block	area of contact with ground
<b>A</b>	decreases	decreases
<b>B</b>	decreases	increases
<b>C</b>	increases	decreases
<b>D</b>	increases	increases

32 A student applies a force to an object, causing the object to move in the same direction as the force.

She measures the size of the force and the distance moved by the object.

Which quantity can she now calculate?

- A** the acceleration of the object
- B** the power she produces
- C** the speed of the object
- D** the work done on the object

33 Cold water evaporates as molecules leave it.

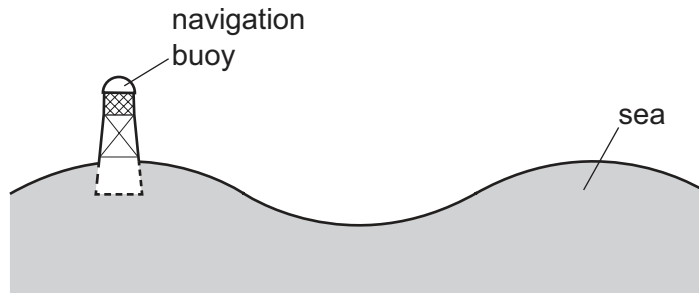
Which molecules leave the water and from which part of the water do they leave?

	molecules that leave the water	where they leave from
<b>A</b>	least energetic	the surface only
<b>B</b>	least energetic	throughout the water
<b>C</b>	most energetic	the surface only
<b>D</b>	most energetic	throughout the water

34 Which material is a good conductor of heat?

- A** copper
- B** glass
- C** plastic
- D** wood

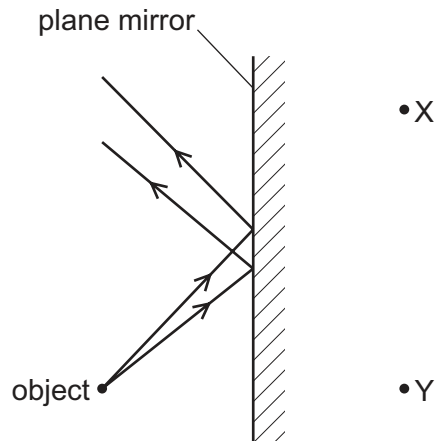
- 35 A navigation buoy floating on the sea oscillates up and down as a wave passes.



In 2.0 minutes, 6.0 wavelengths pass the buoy.

What is the frequency of the waves?

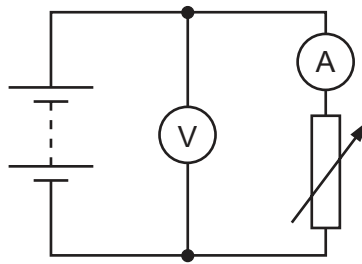
- A 0.050 Hz      B 0.33 Hz      C 3.0 Hz      D 20 Hz
- 36 The diagram shows rays of light from an object being reflected by a plane mirror.



At which labelled point is the image formed, and is the image real or virtual?

	image	real or virtual
<b>A</b>	at X	real
<b>B</b>	at X	virtual
<b>C</b>	at Y	real
<b>D</b>	at Y	virtual

- 37 The diagram represents a circuit that includes a battery, an ammeter, a voltmeter and a variable resistor.

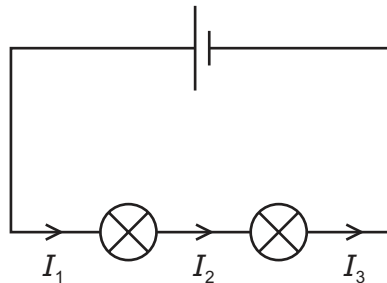


What happens to the readings on the meters as the resistance of the variable resistor is increased?

	ammeter reading	voltmeter reading
<b>A</b>	decreases	decreases
<b>B</b>	decreases	stays constant
<b>C</b>	increases	decreases
<b>D</b>	increases	stays constant

- 38 Two lamps are connected in the circuit shown.

The currents at three points are labelled  $I_1$ ,  $I_2$  and  $I_3$ .

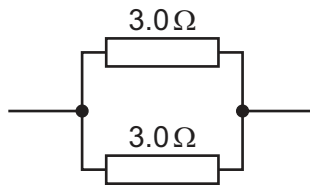


How are  $I_1$ ,  $I_2$  and  $I_3$  related? Use the key to help you.

key  
 < less than  
 > greater than  
 = equal to

- A**  $I_1 < I_2 < I_3$   
**B**  $I_1 = I_2 = I_3$   
**C**  $I_1 > I_2$  and  $I_1 = I_3$   
**D**  $I_1 > I_2 > I_3$

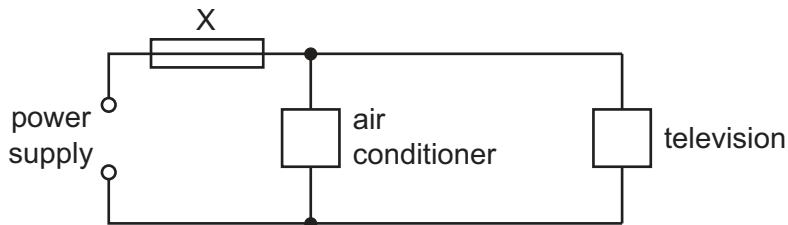
39 Two resistors are connected as shown.



What is the combined resistance of the two resistors?

- A less than  $3.0\ \Omega$
- B  $3.0\ \Omega$
- C  $6.0\ \Omega$
- D  $9.0\ \Omega$

40 An air conditioner and a television are both connected to the same electrical circuit.



The current in the air conditioner is  $9.0\ \text{A}$  and the current in the television is  $2.0\ \text{A}$ .

Several different fuses are available.

Which fuse should be connected at X?

- A 1 A
- B 3 A
- C 7 A
- D 13 A

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

To avoid the issue of disclosure of answer-related information to candidates, all copyright acknowledgements are reproduced online in the Cambridge Assessment International Education Copyright Acknowledgements Booklet. This is produced for each series of examinations and is freely available to download at [www.cambridgeinternational.org](http://www.cambridgeinternational.org) after the live examination series.

Cambridge Assessment International Education is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of the University of Cambridge Local Examinations Syndicate (UCLES), which itself is a department of the University of Cambridge.

The Periodic Table of Elements

		Group															
I	II	III	IV	V	VI	VII	VIII										
3 <b>Li</b> lithium 7	4 <b>Be</b> beryllium 9	5 <b>B</b> boron 11	6 <b>C</b> carbon 12	7 <b>N</b> nitrogen 14	8 <b>O</b> oxygen 16	9 <b>F</b> fluorine 19	10 <b>Ne</b> neon 20					2 <b>He</b> helium 4					
11 <b>Na</b> sodium 23	12 <b>Mg</b> magnesium 24	13 <b>Al</b> aluminium 27	14 <b>Si</b> silicon 28	15 <b>P</b> phosphorus 31	16 <b>S</b> sulfur 32	17 <b>Cl</b> chlorine 35.5	18 <b>Ar</b> argon 40										
19 <b>K</b> potassium 39	20 <b>Ca</b> calcium 40	21 <b>Sc</b> scandium 45	22 <b>Ti</b> titanium 48	23 <b>V</b> vanadium 51	24 <b>Cr</b> chromium 52	25 <b>Mn</b> manganese 55	26 <b>Fe</b> iron 56	27 <b>Co</b> cobalt 59	28 <b>Ni</b> nickel 59	29 <b>Cu</b> copper 64	30 <b>Zn</b> zinc 65	31 <b>Ga</b> gallium 70	32 <b>Ge</b> germanium 73	33 <b>As</b> arsenic 75	34 <b>Se</b> selenium 79	35 <b>Br</b> bromine 80	36 <b>Kr</b> krypton 84
37 <b>Rb</b> rubidium 85	38 <b>Sr</b> strontium 88	39 <b>Y</b> yttrium 89	40 <b>Zr</b> zirconium 91	41 <b>Nb</b> niobium 93	42 <b>Mo</b> molybdenum 96	43 <b>Tc</b> technetium —	44 <b>Ru</b> ruthenium 101	45 <b>Rh</b> rhodium 103	46 <b>Pd</b> palladium 106	47 <b>Ag</b> silver 108	48 <b>Cd</b> cadmium 112	49 <b>In</b> indium 115	50 <b>Sn</b> tin 119	51 <b>Sb</b> antimony 122	52 <b>Te</b> tellurium 128	53 <b>I</b> iodine 127	54 <b>Xe</b> xenon 131
55 <b>Cs</b> caesium 133	56 <b>Ba</b> barium 137	57–71 lanthanoids	72 <b>Hf</b> hafnium 178	73 <b>Ta</b> tantalum 181	74 <b>W</b> tungsten 184	75 <b>Re</b> rhenium 186	76 <b>Os</b> osmium 190	77 <b>Ir</b> iridium 192	78 <b>Pt</b> platinum 195	79 <b>Au</b> gold 197	80 <b>Hg</b> mercury 201	81 <b>Tl</b> thallium 204	82 <b>Pb</b> lead 207	83 <b>Bi</b> bismuth 209	84 <b>Po</b> polonium —	85 <b>At</b> astatine —	86 <b>Rn</b> radon —
87 <b>Fr</b> francium —	88 <b>Ra</b> radium —	89–103 actinoids	104 <b>Rf</b> rutherfordium —	105 <b>Db</b> dubnium —	106 <b>Sg</b> seaborgium —	107 <b>Bh</b> bohrium —	108 <b>Hs</b> hassium —	109 <b>Mt</b> meitnerium —	110 <b>Ds</b> darmstadtium —	111 <b>Rg</b> roentgenium —	112 <b>Cn</b> copernicium —	114 <b>Fl</b> flerovium —	116 <b>Lv</b> livermorium —	—	—	—	—

**Key**

atomic number
atomic symbol
name
relative atomic mass

lanthanoids	57 <b>La</b> lanthanum 139	58 <b>Ce</b> cerium 140	59 <b>Pr</b> praseodymium 141	60 <b>Nd</b> neodymium 144	61 <b>Pm</b> promethium —	62 <b>Sm</b> samarium 150	63 <b>Eu</b> europium 152	64 <b>Gd</b> gadolinium 157	65 <b>Tb</b> terbium 159	66 <b>Dy</b> dysprosium 163	67 <b>Ho</b> holmium 165	68 <b>Er</b> erbium 167	69 <b>Tm</b> thulium 169	70 <b>Yb</b> ytterbium 173	71 <b>Lu</b> lutetium 175
actinoids	89 <b>Ac</b> actinium —	90 <b>Th</b> thorium 232	91 <b>Pa</b> protactinium 231	92 <b>U</b> uranium 238	93 <b>Np</b> neptunium —	94 <b>Pu</b> plutonium —	95 <b>Am</b> americium —	96 <b>Cm</b> curium —	97 <b>Bk</b> berkelium —	98 <b>Cf</b> californium —	99 <b>Es</b> einsteinium —	100 <b>Fm</b> fermium —	101 <b>Md</b> mendelevium —	102 <b>No</b> nobelium —	103 <b>Lr</b> lawrencium —

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).